**Chemistry**

**Fill in the blanks.**

1. The fuel used in Rockets is …………………………….
2. Natural Gas b. Petroleum c. Acetylene d. Liquid Hydrogen
3. …………………… is the number of moles of solute dissolved per liter of solution.
4. Molality b. Normality c. Molarity d. Solvent
5. The chemical formula of Magnetic oxide ………………………
6. Fe2O3 b. Fe O c. Fe3O4 d. Fe2O3.2H2O
7. If 8 g of NaOH is dissolved in 250 ml of water. What will be the Molarity of Solution?
8. 0.32 b. 3.125 c. 8 d. 0.8
9. Bleaching powder contains ………………………..
10. Fluorine b. Chlorine c. Bromine d. Lodine
11. Different physical forms of the same element are called ………………..
12. Allotropes b. Isomers c. Isobars d. Isotopes
13. The symbol of Silicon is …………………..
14. Se b. S c. Si d. Sn
15. Galena is an ore of ……………………………
16. Sulphur b. Iron c. Aluminum d. Copper
17. OH is a functional group of ……………………………..
18. Alkane b. Alkene c. Alcohol d. Alkyne
19. The Radioactive element of 7th group of periodic table is ……………………..
20. Chlorine b. Bromine c. Fluorine d. Astatine
21. Chemical formula of Hydrogen peroxide is …………………………..
22. HO2 b. H2O c. H3O d. H2O2
23. An alloy of Copper and Zinc is ………………………………….
24. Steel b. Brass c. Nichrome d. Bronze
25. Brine is a solution which contains ……………………………..
26. NH3 b. CaCO3 c. NaCl d. Na2CO3
27. Oxygen is present in air by Volume ……………………….
28. 78% b. 21% c. 29% d. 22%
29. Temporary hardness of water is due to ……………………..
30. CaCl2 b. CaSO4 c. Ca(HCO3)2 d. CaCO3
31. Chemical formula of Iron pyrite is …………………………….
32. Fe2O3 b. Fe3O4 c. FeS2 d. Cu2S.Fe2S3
33. Chemical formula of Sodium thiosulphate is ……………………
34. Na2S2O7 b. Na2SO3 c. Na2S2O3 d. Na2S3O3
35. …………………………. Is malleable and ductile
36. Carbon b. Iron c. Sulphur d. Phosphorous
37. Teflon is a ……………………..
38. Soap b. Detergent c. Plastic d. Varnish
39. The Crystalline form of Silica is ……………………………
40. Bauxite b. Limonite c. Sedrite d. Quartz
41. Blood plasma contains ……………………… % of water.
42. 70 b. 80 c. 65 d. 90
43. Plastic Sulphur is stable at …………………………..
44. 71CO b. 96 CO  c. 119 CO d. 141 CO
45. The general formula for alkene is ………………………..
46. CnH2n+2 b. CnH2n c. CnH2n-2 d. CnH2n+1
47. KO2 is an example of ………………………………
48. Acidic oxide b. Per oxide c. Super oxide d. amphoteric oxide
49. Number of moles present in 198 g of water are ………………………..
    1. 19.8 b. 198 c. 11 d. 12.375